



Name: M. Umar Farooq | Quiz Subject:

Chemistry

Time Remaining: 44/45 (Minutes)

Q.1

Test 4 Liquids

CHEMISTRY NMDCAT

Which correctly represent the increasing order of boiling points of hydrides of IV-A group?

- a. $\text{CH}_4 < \text{GeH}_4 < \text{SiH}_4 < \text{SnH}_4$
- b. $\text{SiH}_4 > \text{CH}_4 < \text{GeH}_4 < \text{SnH}_4$
- c. $\text{CH}_4 < \text{SiH}_4 < \text{GeH}_4 < \text{SnH}_4$
- d. $\text{GeH}_4 < \text{SnH}_4 < \text{H}_4 < \text{SiH}_4$

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Correct Answer:

☐ A ☐ B ☐ C ☐ D

Next

Time Remaining: 44/45 (Minutes)

Q.2

Test 4 Liquids

CHEMISTRY NMDCAT

Ethyl Alcohol is soluble in water due to:

- a. Dipole-dipole forces
- b. Dipole-induced dipole forces
- c. Instantaneous dipole-induced dipole
- d. Hydrogen bonding

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Correct Answer:

☒ A ☐ B ☐ C ☐ D

Next

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Time Remaining: 44/45 (Minutes)

Q.3

Test 4 Liquids

CHEMISTRY NMDCAT

What about the solubility of hydrocarbons in H_2O ?

- a. Readily soluble
- b. Partially soluble
- c. Insoluble
- d. Soluble at high temperature

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Correct Answer:

☒ A ☐ B ☐ C ☐ D

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Time Remaining: 43/45 (Minutes)

Q.4

Test 4 Liquids

CHEMISTRY NMDCAT

With the increase in Vander Waal's forces the non-ideality in a gas:

- a. Increases
- b. Decreases
- c. Both a & b
- d. have no effect

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Correct Answer:

☒ A ☐ B ☐ C ☐ D

Next

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Time Remaining: 43/45 (Minutes)

Q.5

Test 4 Liquids

CHEMISTRY NMDCAT

Which one of the following intermolecular forces has minimum strength?

- a. H-bonding
- b. Dipole-dipole interaction
- c. London forces
- d. Dipole-induced dipole forces

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Correct Answer:

☒ A ☐ B ☐ C ☐ D

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Time Remaining: 43/45 (Minutes)

Q.6

Test 4 Liquids

CHEMISTRY NMDCAT

Long chains of Amino acids are coiled about one another into a spiral by:

- a. Covalent bond
- b. Ionic bonds
- c. Hydrogen bonds
- d. Van der Waal's forces

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Correct Answer:

☒ A ☐ B ☐ C ☐ D

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Time Remaining: 43/45 (Minutes)

Q.7

Test 4 Liquids

CHEMISTRY NMDCAT

In which compound hydrogen bonding is not present?

a. H_2O

b. NH_3

c. CH_4

d. C_2H_5OH

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Correct Answer:



A



B



C



D

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Time Remaining: 43/45 (Minutes)

Q.8

Test 4 Liquids

CHEMISTRY NMDCAT

The deciding criteria for determination of melting and boiling points in case of group-IV members compounds is?

- a. Hydrogen bonding in compound
- b. Intermolecular forces in compound
- c. Chemical bonds in compound
- d. Molar mass of compound

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Correct Answer:

☒ A ☐ B ☐ C ☐ D

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Time Remaining: 43/45 (Minutes)

Q.9

Test 4 Liquids

CHEMISTRY NMDCAT

Which is the correct order of stability?

- a. hydrogen bonding > dipole-dipole > London forces
- b. dipole-dipole > London forces > hydrogen bonding
- c. London forces > hydrogen bonding > dipole-dipole
- d. hydrogen bonding > London forces > dipole-dipole

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Correct Answer:

☒ A ☐ B ☐ C ☐ D

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Time Remaining: 43/45 (Minutes)

Q.10

Test 4 Liquids

CHEMISTRY NMDCAT

Which does not affect vapour pressure?

- a. Nature of liquid
- b. Intermolecular forces
- c. temp
- d. none of these

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Correct Answer:

☒ A ☐ B ☐ C ☐ D

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Time Remaining 43/45 (Minutes)

Q.11

Test 4 Liquids

CHEMISTRY NMDCAT

Heat of vaporization increased by

- a. ionic bonding b. Covalent bond
c. Vander wall's d. H.B

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Correct Answer:



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Chemistry

Time Remaining 42/45 (Minutes)

Q.111

Test 4 Liquids

CHEMISTRY NMDCAT

_____ is the measure of average K.E of molecules of liquid.

- a. velocity b. heat
c. temp d. speed

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Correct Answer:



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Chemistry

Time Remaining 42/45 (Minutes)

Q.11

Test 4 Liquids

CHEMISTRY NMDCAT

The hydrogen bonding is absent in:

- a. Ice
- b. CHCl_3 & CH_3COCH_3
- c. Steam
- d. DNA

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Correct Answer:



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Time Remaining 42/45 (Minutes)

QUIZ

Test 4 Liquids

CHEMISTRY NMDCAT

Which one is max. soluble in H_2O ?

- a. CH_3OH b. C_2H_5OH
c. C_3H_7OH d. $C_6H_5CH_2OH$

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Correct Answer:



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Time Remaining 42/45 (Minutes)

Q.19

Test 4 Liquids

CHEMISTRY NMDCAT

Least vapour pressure at 25°C:

- | | |
|----------------|---------------------|
| a. Ethanol | b. Acetone |
| c. Acetic acid | d. H ₂ O |

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Correct Answer:



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Chemistry

Time Remaining 41/45 (Minutes)

Q.10

Test 4 Liquids

CHEMISTRY NMDCAT

Ethylene glycol has less V.P than ethanol be:

- a. Molecular mass
- b. Size of molecule
- c. H-bond
- d. Ethylene less volatile than C_2H_5OH

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Correct Answer:



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Time Remaining 41/45 (Minutes)

Q.17

Test 4 Liquids

CHEMISTRY NMDCAT

Which one is most important parameter which control V.P?

- a. Surface area
- b. Forces
- c. Temperature
- d. Nature of liquid

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Correct Answer:



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Time Remaining 41/45 (Minutes)

Q.117

Test 4 Liquids

CHEMISTRY NMDCAT

Pressure cooker cooks food faster:

- a. B.P decrease
- b. More heat is absorbed
- c. External "P" increase artificially
- d. V.P of H_2O increase

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Correct Answer:



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Time Remaining 41/45 (Minutes)

Q.43

Test 4 Liquids

CHEMISTRY NMDCAT

Highest Heat of vaporization is for?

- a. F_2 b. Cl_2
c. Br_2 d. I_2

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Correct Answer:



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Time Remaining 41/45 (Minutes)

Q.25

Test 4 Liquids

CHEMISTRY NMDCAT

Strength of dipole dipole forces depend on

- (a) electro negativity difference
- (b) distance between two molecules
- (c) both a & b
- (d) number of atoms in a molecule

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Correct Answer:



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Chemistry

Time Remaining 41/45 (Minutes)

Q11

Test 4 Liquids

CHEMISTRY NMDCAT

Which of the following has exceptionally low acidic strength

- (a) HI
- (b) HBr
- (c) HF
- (d) HCl

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Correct Answer:



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Time Remaining 41/45 (Minutes)

Q.11

Test 4 Liquids

CHEMISTRY NMDCAT

Forces of attraction which may be present between all kinds of atoms and molecule are

- (a) Van der Waal's forces
- (b) intra molecular
- (c) london dispersion forces
- (d) dipole induced dipole force

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Correct Answer:



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Chemistry

Time Remaining 40/45 (Minutes)

Q.10

Test 4 Liquids

CHEMISTRY NMDCAT

All elements of the zero group in the periodic table possess

- (a) Ion dipole forces
- (b) London forces
- (c) Hydrogen bonding
- (d) dipole dipole forces

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Correct Answer:



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Chemistry

Time Remaining 40/45 (Minutes)

Q.14

Test 4 Liquids

CHEMISTRY NMDCAT

Hydrogen bonding is maximum in

- (a) ethanol
- (b) benzene
- (c) diethyl ether
- (d) water

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Correct Answer:



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Chemistry

Time Remaining 40/45 (Minutes)

Q.30

Test 4 Liquids

CHEMISTRY NMDCAT

The weakest intermolecular forces present in a liquid may be

- (a) dipole induced dipole forces
- (b) dipole dipole forces
- (c) london dispersion forces
- (d) hydrogen bonding

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Correct Answer:



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Chemistry

Time Remaining 40/45 (Minutes)

Q.30

Test 4 Liquids

CHEMISTRY NMDCAT

When liquid water changes to ice its volume expands by

- (a) 100% (b) 5%
(c) 9% (d) 4%

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Correct Answer:



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Chemistry

Time Remaining 39/45 (Minutes)

Q.11

Test 4 Liquids

CHEMISTRY NMDCAT

Down the group polarizability generally

- (a) increases (b) decreases
(c) remains same (d) none of these

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Correct Answer:



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Chemistry

Time Remaining 39/45 (Minutes)

Q.18

Test 4 Liquids

CHEMISTRY NMDCAT

Which one of the following forces is the strongest

- (a) Hydrogen bonding
- (b) Dipole dipole attraction
- (c) London dispersion forces
- (d) Debye forces

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Correct Answer:



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Chemistry

Time Remaining 39/45 (Minutes)

02:29

Test 4 Liquids

CHEMISTRY NMDCAT

Water at 100°C has

- (a) greater energy than steam
- (b) same energy as steam
- (c) less energy than steam
- (d) no energy

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Correct Answer:



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Chemistry

Time Remaining 39/45 (Minutes)

Q35

Test 4 Liquids

CHEMISTRY NMDCAT

Conversion of vapours back into their liquid state is called

- (a) crystallization
- (b) distillation
- (c) vapourization
- (d) condensation

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Correct Answer:



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Chemistry

Time Remaining 38/45 (Minutes)

Q31

Test 4 Liquids

CHEMISTRY NMDCAT

If we provide a very large amount of heat to liquid its boiling point will

- (a) Increase
- (b) Remain constant
- (c) Decrease
- (d) There will be no boiling

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Correct Answer:



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Time Remaining 37/45 (Minutes)

Q.31

Test 4 Liquids

CHEMISTRY NMDCAT

Vapour pressure is not affected by

- (a) nature of liquid
- (b) temperature
- (c) intermolecular forces
- (d) amount of liquid

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Correct Answer:



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Chemistry

Time Remaining 37/45 (Minutes)

Q33

Test 4 Liquids

CHEMISTRY NMDCAT

When the external pressure is 23.7 torr, boiling point of water is

- (a) 25°C (b) 100°C
(c) 98°C (d) 200°C

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Correct Answer:



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Time Remaining 37/45 (Minutes)

Q.34

Test 4 Liquids

CHEMISTRY NMDCAT

The boiling point of the halogens

- (a) increases down the group
- (b) decreases down the group
- (c) remains constant
- (d) difficult to predict

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Correct Answer:



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Chemistry

Time Remaining 37:45 (Minutes)

Q.39

Test 4 Liquids

CHEMISTRY NMDCAT

The distillation of a solution under reduced pressure is called

- (a) fractional distillation
- (b) destructive distillation
- (c) dry distillation
- (d) vacuum distillation

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Correct Answer:



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Time Remaining 36/45 (Minutes)

Q38

Test 4 Liquids

CHEMISTRY NMDCAT

The boiling point of a compound is mostly raised by

- (a) intermolecular forces
- (b) intramolecular forces
- (c) both a & b
- (d) none of these

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Correct Answer:



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Chemistry

Time Remaining 36/45 (Minutes)

QJF

Test 4 Liquids

CHEMISTRY NMDCAT

The value of vapour pressure of water at its boiling point in Karachi and Murree is

- (a) greater at Murree and lower in Karachi
- (b) same
- (c) lower at Murree and higher in Karachi
- (d) none of these

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Correct Answer:



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Time Remaining 36/45 (Minutes)

Q.38

Test 4 Liquids

CHEMISTRY NMDCAT

Which of the following sequence is correct?

- (a) $\Delta H_s > \Delta H_v > \Delta H_f$
- (b) $\Delta H_v > \Delta H_s > \Delta H_f$
- (c) $\Delta H_f > \Delta H_s > \Delta H_v$
- (d) all have same enthalpy

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Correct Answer:



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Time Remaining 36/45 (Minutes)

Q39

Test 4 Liquids

CHEMISTRY NMDCAT

Which one of the following have lowest vapour pressure at 25° C

- (a) water
- (b) ethyl alcohol
- (c) acetone
- (d) diethyl ether

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Correct Answer:



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Chemistry

Time Remaining 35/45 (Minutes)

Q.45

Test 4 Liquids

CHEMISTRY NMDCAT

Molar heat of vapourization of water is

- (a) 40.7 kJ mol^{-1} (b) 50.9 kJ mol^{-1}
(c) $40.7 \text{ kcal mol}^{-1}$ (d) $50.7 \text{ kcal mol}^{-1}$

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Correct Answer:



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Q. 1

Which correctly represent the increasing order of boiling points of hydrides of IV-A group?

- a. $\text{CH}_4 < \text{GeH}_4 < \text{SiH}_4 < \text{SnH}_4$
- b. $\text{SiH}_4 > \text{CH}_4 < \text{GeH}_4 < \text{SnH}_4$
- c. $\text{CH}_4 < \text{SiH}_4 < \text{GeH}_4 < \text{SnH}_4$
- d. $\text{GeH}_4 < \text{SnH}_4 < \text{H}_4 < \text{SiH}_4$

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Q. 2

Ethyl Alcohol is soluble in water due to:

- a. Dipole-dipole forces
- b. Dipole-induced dipole forces
- c. Instantaneous dipole-induced dipole
- d. **Hydrogen bonding**

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Q. 3

What about the solubility of hydrocarbons in H_2O ?

- a. Readily soluble
- b. Partially soluble
- c. **Insoluble**
- d. Soluble at high temperature

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Q. 4

With the increase in Vander Waal's forces the non-ideality in a gas:

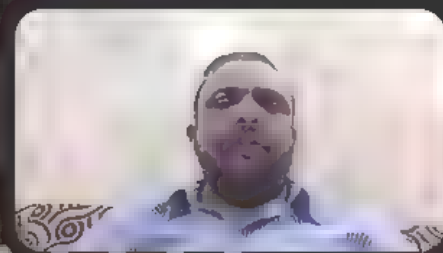
- a. Increases
- b. Decreases
- c. Both a & b
- d. have no effect

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Q. 5

Which one of the following intermolecular forces has minimum strength?

- a. H-bonding
- b. Dipole-dipole interaction
- c. **London forces**
- d. Dipole-induced dipole forces

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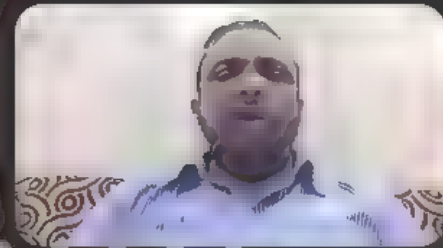
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Q. 6

Long chains of Amino acids are coiled about one another into a spiral by:

- a. Covalent bond
- b. Ionic bonds
- c. **Hydrogen bonds**
- d. Van der Waal's forces

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Q. 6

Long chains of Amino acids are coiled about one another into a spiral by:

- a. Covalent bond
- b. Ionic bonds
- c. **Hydrogen bonds**
- d. Van der Waal's forces

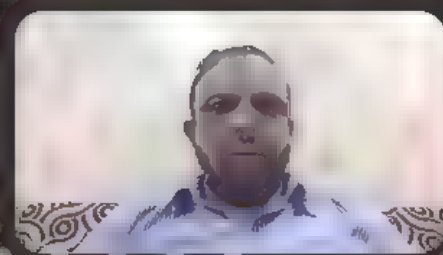
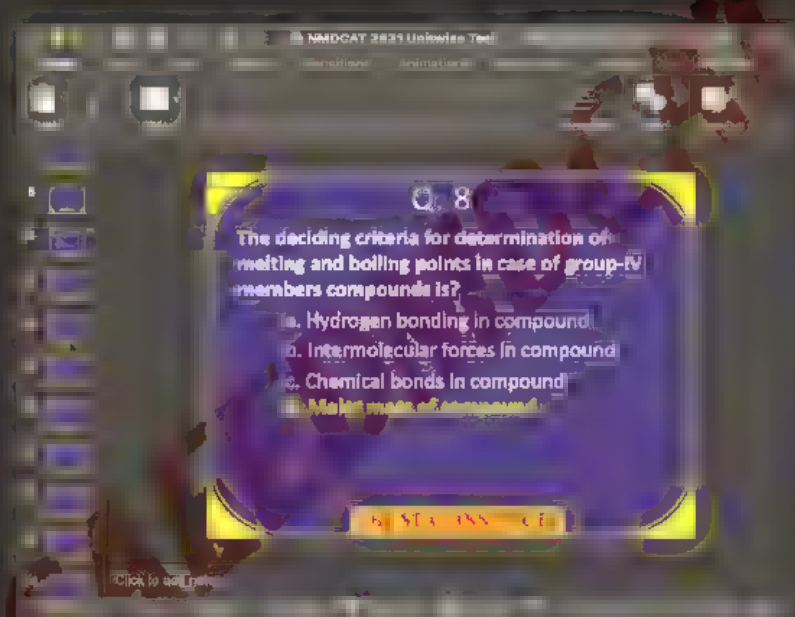
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Q. 7

In which compound hydrogen bonding is not present?



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Q. 8

The deciding criteria for determination of melting and boiling points in case of group-IV members compounds is?

- a. Hydrogen bonding in compound
- b. Intermolecular forces in compound
- c. Chemical bonds in compound
- d. Molar mass of compound**

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Q. 9

Which is the correct order of stability?

a. hydrogen bonding > dipole-dipole > London forces

b. dipole-dipole > London forces > hydrogen bonding

c. London forces > hydrogen bonding > dipole-dipole

d. hydrogen bonding > London forces > dipole-dipole

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Q. 10

Which does not affect vapour pressure?

- a. Nature of liquid
- b. Intermolecular forces
- c. temp
- d. none of these

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Q.11

Heat of vaporization increased by

- a. ionic bonding
- b. Covalent bond
- c. Vander wall's
- d. H.B

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Q. 12

_____ is the measure of average K.E of molecules of liquid.

a. velocity

b. heat

c. temp.

d. speed

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Q. 13

The hydrogen bonding is absent in:

- a. Ice b. CHCl_3 & CH_3COCH_3
c. Steam d. DNA

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Q. 14

Which one is max. soluble in H_2O ?

a. CH_3OH

b. C_2H_5OH

c. C_3H_7OH

d. $C_6H_5CH_2OH$

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Q. 15

Least vapour pressure at 25°C:

- a. Ethanol
- b. Acetone
- c. Acetic acid
- d. H_2O

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Q. 16

Ethylene glycol has less V.P than ethanol be:

- a. Molecular mass
- b. Size of molecule
- c. H-bond
- d. Ethylene less volatile than C_2H_5OH

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Q. 17

Which one is most important parameter which control V.P?

- ☐ a. Surface area
- ☐ b. Forces
- ☒ c. Temperature
- ☐ d. Nature of liquid

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Q. 18

Pressure cooker cooks food faster:

- a. B.P decrease
- b. More heat is absorbed
- c. External P increase artificially
- d. V.P of H_2O increase

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Q. 19

Highest Heat of vaporization is for?

a. F_2 b. Cl_2 c. Br_2 ☒ d. I_2

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Q. 20

Strength of dipole dipole forces depend on

- (a) electro negativity difference
- (b) distance between two molecules
- (c) both a & b
- (d) number of atoms in a molecule

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Q. 21

Which of the following has exceptionally low acidic strength

- (a) HI
- (b) HBr
- (c) HF
- (d) HCl

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Q. 22

Forces of attraction which may be present between all kinds of atoms and molecule are

- (a) Van der Waal's forces
- (b) intra molecular
- (c) London dispersion forces
- (d) dipole induced dipole force

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Q. 23

All elements of the zero group in the periodic table possess

- (a) Ion dipole forces
- (b) London force**
- (c) Hydrogen bonding
- (d) dipole dipole forces

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Q. 24

Hydrogen bonding is maximum in

- (a) ethanol
- (b) benzene
- (c) diethyl ether
- (d) water

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Q. 25

The weakest intermolecular forces present in a liquid may be

- (a) dipole induced dipole forces
- (b) dipole dipole forces
- (c) london dispersion forces
- (d) hydrogen bonding

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Q. 26

When liquid water changes to ice its volume expands by

(a) 100%

(b) 5%

(c) 9%

(d) 4%

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Q. 27

Down the group polarizability generally

☒ (a) increase☐ (b) decrease☐ (c) remains same☐ (d) none of these

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Q. 28

Which one of the following forces is the strongest

- (a) Hydrogen bonding
- (b) Dipole dipole attraction
- (c) London dispersion forces
- (d) Debye forces

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Q. 29

Water at 100°C has

- (a) greater energy than steam
- (b) same energy as steam
- (c) less energy than steam
- (d) no energy

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Q. 30

Conversion of vapours back into their liquid state is called

- (a) crystallization
- (b) distillation
- (c) vapourization
- (d) condensation

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Q. 31

If we provide a very large amount of heat to liquid its boiling point will

- (a) Increase
- (b) Remain constant
- (c) Decrease
- (d) There will be no boiling

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Q. 32

Vapour pressure is not affected by

- (a) nature of liquid
- (b) temperature
- (c) intermolecular forces
- (d) amount of liquid

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Q. 33

When the external pressure is 23.7 torr, boiling point of water is

- (a) 25°C
- (b) 100°C
- (c) 98°C
- (d) 200°C

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Q. 34

The boiling point of the halogens

- (a) increases down the group
- (b) decreases down the group
- (c) remains constant
- (d) difficult to predict

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Q. 35

The distillation of a solution under reduced pressure is called

- (a) fractional distillation
- (b) destructive distillation
- (c) dry distillation
- (d) vacuum distillation

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Q. 36

The boiling point of a compound is mostly raised by

- (a) intermolecular forces
- (b) intramolecular forces
- (c) both a & b
- (d) none of these

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Q. 37

The value of vapour pressure of water at its boiling point in Karachi and Murree is

- (a) greater at Murree and lower in Karachi
- (b) same
- (c) lower at Murree and higher in Karachi
- (d) none of these

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Q. 38

Which of the following sequence is correct?

- (a) $\Delta H_s > \Delta H_v > \Delta H_f$
- (b) $\Delta H_v > \Delta H_s > \Delta H_f$
- (c) $\Delta H_f > \Delta H_s > \Delta H_v$
- (d) all have same enthalpy

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Q. 39

Which one of the following have lowest vapour pressure at 25° C

- (a) water
- (b) ethyl alcohol
- (c) acetone
- (d) diethyl ether

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Q. 40

Molar heat of vapourization of water is

- (a) 40.7 kJ mol^{-1} (b) 50.9 kJ mol^{-1}
(c) $40.7 \text{ k cal mol}^{-1}$ (d) $50.7 \text{ k cal mol}^{-1}$

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